A. **Academic Division:** Health Sciences

B. **Discipline:** Science

C. **Course Number and Title:** CHEM 1220 Chemistry II

D. **Course Coordinator:**
   - **Assistant Dean:** Melinda S. Roepke, MSN, RN

**Instructor Information:**
- **Name:** Click here to enter text.
- **Office Location:** Click here to enter text.
- **Office Hours:** Click here to enter text.
- **Phone Number:** Click here to enter text.
- **E-Mail Address:** Click here to enter text.

E. **Credit Hours:** 5
   - Lecture: 4 hours
   - Laboratory: 3 hours

F. **Prerequisites:** CHEM1210

G. **Syllabus Effective Date:** Fall, 2017

H. **Textbook(s) Title:**
   - *Chemistry: The Central Science*
     - Author(s): Brown, LeMay, Bursten, Murphy
     - Copyright Year: 2018
     - Edition: 14th
     - ISBN: 9780134809663

   **Optional Textbooks:** *Student’s Guide to Accompany Chemistry: The Central Science*
     - Author(s): Hill
     - Copyright Year: 2012
     - Edition: 12
     - ISBN #: 978-0321704580

I. **Workbook(s) and/or Lab Manual:**
     - Author(s):
     - Copyright Year: Produced for Ohio State University
     - Edition:
     - ISBN #: Customized Lab Handouts
J. **Course Description:** This is the second semester of chemistry for science majors or pre-professional students. A quantitative introduction to intermolecular forces, phase changes, colligative properties, chemical kinetics, chemical equilibrium, acid-base equilibria, thermodynamic, electrochemistry, organic chemistry and nuclear chemistry. Student will be exposed to applications of chemistry in society. (TAG # OSC009; If combined with CHEM1210 TAG # OSC023)

K. **College-Wide Learning Outcomes**

<table>
<thead>
<tr>
<th>College-Wide Learning Outcomes</th>
<th>Assessments - - How it is met &amp; When it is met</th>
</tr>
</thead>
<tbody>
<tr>
<td>Communication – Written</td>
<td></td>
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<tr>
<td>Communication – Speech</td>
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<tr>
<td>Intercultural Knowledge and Competence</td>
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<td>Critical Thinking</td>
<td></td>
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<td>Information Literacy</td>
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<tr>
<td>Quantitative Literacy</td>
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</tbody>
</table>

L. **Course Outcomes and Assessment Methods:**

Upon successful completion of this course, the student shall:

<table>
<thead>
<tr>
<th>Outcomes</th>
<th>Assessments – How it is met &amp; When it is met</th>
</tr>
</thead>
</table>
| 1. Demonstrate intermolecular forces and phase diagrams. | Quiz - 2nd week  
Written lab report  
Midterm Exam – 4th week  
Final – 15th week |
| 2. Describe the solution process and demonstrate colligative properties. | Quiz – 3rd week  
Written Lab Report  
Mid-Term Exam – 4th week  
Final – 15th week |
| 3. Explain chemical kinetics | Written Lab Report  
Mid-Term Exam – 4th week  
Final – 15th week |
| 4. Demonstrate chemical equilibrium | Quiz – 5th week  
Written Lab Report  
Mid-Term Exam-8th week  
Final Exam – 15th week |
| 5. Explain and demonstrate acid-base equilibria and buffers. | Quiz-6th, 7th week  
Written Lab Reports  
Mid-Term Exam-8th week  
Final Exam – 15th week m – 8th week |
| 6. Describe chemical thermodynamics, entropy and free energy. | Quiz-9th week  
Mid-Term Exam- 12th week  
Final Exam – 15th week |
| 7. Explain and demonstrate electrochemistry. | Quiz- 10th & 11th week  
Written Lab Report  
Mid-Term Exam-12th week |
| 8. Define organic chemistry. | Quiz- 13th week  
Formal Written Lab Report  
Mid-Term Exam- 12th week |
| 9. Explain transition metals & coordination chemistry. | Quiz – 14th week  
Final Exam – 15th week |
| 10. Describe nuclear chemistry. | Quiz- 14th week  
Final Exam 15th week |
M. Topical Timeline (Subject to Change):

**Lecture Topics**

1. Liquids and Intermolecular Forces: A molecular comparison of gases, liquids and solids, Intermolecular forces Select properties of liquids, Phase changes, Vapor pressure, Phase diagrams.
4. Chemical Kinetics: Reaction rates, Concentration and rate laws, Concentration vs. time, Temperature and rate, Reaction mechanisms.
5. Chemical Equilibrium: The concept of equilibrium, Equilibrium constants, Heterogeneous equilibria, Calculating equilibrium constants (ICE tables), Applications of equilibrium constants, Le Châtelier’s principle.
6. Acid Base Equilibria: Acid-Base Equilibria, Bronsted-Lowry acids and bases, the pH Scale, Strong acids and bases, Weak acids and bases, Relationship between Ka and Kb, pH of salt solutions, Chemical structure and acidity/basicity, Lewis acids and bases.
11. Transition metals and coordination chemistry: The transition metals, coordination complexes, Ligands, Nomenclature, Isomerism, Color and magnetism of transition metal compounds, Crystal field theory.
12. Nuclear Chemistry: Radioactivity, Patterns of nuclear stability, Nuclear transmutations, Rates of radioactive delay, Energy changes in nuclear reactions, Nuclear fission and fusion, Radiation in the environment and living systems.

**Laboratory exercises**

1. Check-in/Intermolecular Forces
2. Freezing Point Depression
3. Rate of an Iodine Clock Reaction
4. Volumetric Analysis: Standardization of NaOH and HCl
5. Equilibrium
6. Investigation of Buffer Systems
7. Some Nonmetals and their Compounds; Hydrolysis of Salts
8. Qualitative Analysis Group I
9. Redox Titration
10. Electrochemistry
11. Spectrophotometric Analysis of Commercial Aspirin
12. Research Experience
N. **Course Assignments:**

1. Homework
2. Formal Lab Reports
3. Quizzes, Mid-Term, Final

O. **Recommended Grading Scale:**

<table>
<thead>
<tr>
<th>NUMERIC</th>
<th>GRADE</th>
<th>POINTS</th>
<th>DEFINITION</th>
</tr>
</thead>
<tbody>
<tr>
<td>93–100</td>
<td>A</td>
<td>4.00</td>
<td>Superior</td>
</tr>
<tr>
<td>90–92</td>
<td>A-</td>
<td>3.67</td>
<td>Superior</td>
</tr>
<tr>
<td>87–89</td>
<td>B+</td>
<td>3.33</td>
<td>Above Average</td>
</tr>
<tr>
<td>83–86</td>
<td>B</td>
<td>3.00</td>
<td>Above Average</td>
</tr>
<tr>
<td>80–82</td>
<td>B-</td>
<td>2.67</td>
<td>Above Average</td>
</tr>
<tr>
<td>77–79</td>
<td>C+</td>
<td>2.33</td>
<td>Average</td>
</tr>
<tr>
<td>73–76</td>
<td>C</td>
<td>2.00</td>
<td>Average</td>
</tr>
<tr>
<td>70–72</td>
<td>C-</td>
<td>1.67</td>
<td>Below Average</td>
</tr>
<tr>
<td>67–69</td>
<td>D+</td>
<td>1.33</td>
<td>Below Average</td>
</tr>
<tr>
<td>63–66</td>
<td>D</td>
<td>1.00</td>
<td>Below Average</td>
</tr>
<tr>
<td>60–62</td>
<td>D-</td>
<td>0.67</td>
<td>Poor</td>
</tr>
<tr>
<td>00–59</td>
<td>F</td>
<td>0.00</td>
<td>Failure</td>
</tr>
</tbody>
</table>

P. **Grading and Testing Guidelines:**

Click here to enter text.

Q. **Examination Policy:**

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R. **Class Attendance and Homework Make-Up Policy:**

Click here to enter text.

S. **Classroom Expectations:**

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T. **College Procedures/Policies:**

**Attendance Requirements:** All students are required to attend all scheduled classes and examinations. Each faculty member has the right to establish regulations regarding attendance that he/she considers necessary for successful study.

Students who do not attend classes may be administratively withdrawn from those classes. However, failure to attend classes does not constitute withdrawal, and students are expected to process a formal withdrawal though the Student Records Office in Kee Hall.

**Student engagement requirements:**

Student engagement is based on the “active pursuit” of learning which can be measured by class attendance, class participation (in class or online), taking required quizzes/examinations, and submission of work assignments or papers. Student engagement consists of a student attending at least 60% of the class sessions (there should be attendance throughout the term) and/or completing 75% of the assignments listed on the syllabus at the midpoint in the term. Exceptions can be made when there is on-going
communication between the student and faculty member. The communication must be documented and the faculty member and student must be in agreement regarding the exception. Students not meeting the expectation will be administratively withdrawn from class. If a student believes he/she was administratively withdrawn in error, he/she may file an appeal. Being administratively withdrawn may have program and financial aid implications.

**Academic Misconduct** is any activity that tends to compromise the academic integrity of the college, or subvert the educational process. Examples of academic misconduct include, but are not limited to:

1. **Violation of course or program rules** as contained in the course syllabus or other information provided to the student; violation of program requirements as established by departments and made available to students.

2. **Plagiarism** including, but not limited to, submitting, without appropriate acknowledgment, any written, visual or oral material that has been copied in whole or in part from the work of others (whether such source is published or not) even if the material is completely paraphrased in one’s own words. This includes another individual’s academic composition, compilation, or other product, or a commercially prepared paper. Plagiarism also includes submitting work in which portions were substantially produced by someone acting as a tutor or editor.

   Such practices constitute plagiarism regardless of motive. Those who deny deceitful intent, claim not to have known that the act constituted plagiarism, or maintain that what they did was inadvertent are nevertheless subject to penalties when plagiarism has been confirmed.

3. **Cheating** and dishonest practices in connection with examinations, papers and projects, including but not limited to using unauthorized notes, study aids or information on an examination; obtaining help from another student during an examination; taking an exam or doing work for another student; providing one’s own work for another student to copy and submit as his/her own; or allowing another student to do one’s work and then submitting the work as one’s own. Also included would be altering a graded work after it has been returned, then submitting the work for re-grading; or submitting identical or similar papers for credit in more than one course without prior permission from the course instructors.

4. **Fabrication** including but not limited to falsifying or inventing any information, data or citation; presenting data that were not gathered in accordance with defined appropriate guidelines, and failing to include an accurate account of the method by which data were collected.

5. **Obtaining an Unfair Advantage** including, but not limited to stealing, reproducing, circulating, or otherwise gaining access to examination materials prior to the time authorized by the instructor; unauthorized collaborating on an academic assignment; taking, hiding or altering resource material; or undertaking any activity with the purpose of creating or obtaining an unfair advantage over another student’s academic work.

6. **Aiding and Abetting Academic Dishonesty** including, but not limited to providing material, information or other assistance to another person with the knowledge that such aid could be used in any of the violations stated above, or providing false information in connection with any inquiry regarding academic integrity.

7. **Alteration of Grades or Marks** including but not limited to, action by the student in an effort to change the earned credit or grade.

In addition, cases of academic dishonesty may involve photocopied materials. Materials used may fall under the Copyright Act. Violations of said Act may subject the user and/or the College to sanctions.
Statement on Disabilities: Any student who requires reasonable accommodations related to a disability should inform the course instructor and the Coordinator of Specialized Services (Room 138 in Kee Hall; phone 419-755-4727).

Students who encounter difficulty in any of their courses are encouraged to visit the Tutoring Resource Center (Room 119 in Fallerius Technical Education Center) for tutoring assistance, and the Student Success Center (Room 136 in Kee Hall) for academic assistance, advising services, referrals for personal counseling and Learning Disability (LD) Testing.

Statement on Withdrawals: As a student, you are expected to attend class. If you are unable or choose not to attend class, or if for whatever reason you are unable to keep up with the requirements of a course, you need to officially drop the class at the Student Records Office. Refund dates and withdrawal dates will vary slightly from term to term. Contact the Student Records Office for applicable dates. Additionally these dates are posted on the academic calendar available on the college’s website, www.ncstatecollege.edu, under the Academics heading on the home page and are available at the Student Records Office in Kee Hall. Students should go to the Student Records Office (Room 142 in Kee Hall) to process their withdrawal from any class.

If you choose to walk away from your class without officially withdrawing from it, the faculty member teaching the class must grade your classroom performance on the material available to him or her. This normally results in an "F" grade. An "F" grade can lower your grade point average considerably depending on the total credits accumulated.